

Three-dimensional model evaluation of the Ozone Depletion Potentials for n-propyl bromide, trichloroethylene and perchloroethylene

D. J. Wuebbles¹, K. O. Patten¹, D. Wang¹, D. Youn², M. Martínez-Avilés^{3,*}, and J. S. Francisco³

¹Department of Atmospheric Sciences, University of Illinois at Urbana-Champaign, Urbana, IL, USA

²School of Earth and Environmental Sciences Seoul National University Seoul, Seoul, South Korea

³Department of Chemistry and Department of Earth and Atmospheric Sciences, Purdue University, West Lafayette, IN, USA

* now at: National Institute of Water and Atmospheric Research Omakau, Central Otago, New Zealand

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Abstract. The existing solvents trichloroethylene (TCE) and perchloroethylene (PCE) and proposed solvent n-propyl bromide (nPB) have atmospheric lifetimes from days to a few months, but contain chlorine or bromine that could affect stratospheric ozone. Several previous studies estimated the Ozone Depletion Potentials (ODPs) for various assumptions of nPB emissions location, but these studies used simplified modeling treatments. The primary purpose of this study is to reevaluate the ODP for n-propyl bromide (nPB) using a current-generation chemistry-transport model of the troposphere and stratosphere. For the first time, ODPs for TCE and PCE are also evaluated in a three-dimensional, global atmospheric chemistry-transport model. Emissions representing industrial use of each compound are incorporated on land surfaces from 30° N to 60° N. The atmospheric chemical lifetime obtained for nPB is 24.7 days, similar to past literature, but the ODP is 0.0049, lower than in our past study of nPB. The derived atmospheric lifetime for TCE is 13.0 days and for PCE is 111 days. The corresponding ODPs are 0.00037 and 0.0050, respectively.

trichloroethylene (TCE, chemical formula C_2HCl_3) and perchloroethylene (PCE, chemical formula C_2Cl_4 , also named tetrachloroethylene). TCE is widely used as an industrial solvent; PCE is commonly used in fabrics dry-cleaning, metal degreasing, as well as a feedstock for chemical manufacture. As discussed in the 2002 World Meteorological Organization international stratospheric ozone assessment (Ko et al., 2003), four modeling studies (Wuebbles et al., 1999, 2001; Bridgeman et al., 2000; Olsen et al., 2000) have previously evaluated Ozone Depletion Potentials (ODPs) for nPB. These studies used the modified concept for ODPs that accounts for the location of emissions (Wuebbles and Ko, 1999) as compared to the original definition developed for longer-lived gases. However, none of the nPB studies evaluated the ODPs using a three-dimensional chemistry-transport model of the troposphere and stratosphere with complete representation of relevant chemistry processes. The purpose of this study is to update the ODPs for nPB using such a model. In addition, ODPs of TCE and PCE are evaluated, though the primary environmental concerns about these compounds relate to their toxicity, as summarized in OEHHA (1999, 2000).

In the earlier nPB studies, Wuebbles et al. (1999) used a zonally-averaged two-dimensional (2-D) model to estimate the ODP. Olsen et al. (2000) applied the low-resolution (8° latitude × 10° longitude × 9 layers) three-dimensional (3-D) UCI tropospheric chemical-transport model to determine the amount of very short-lived (VSL) substances for a range of atmospheric lifetimes, and for various seasons and latitudes of emission, reaching the tropical tropopause. Then by scaling the ODP values from 2-D model results (Wuebbles et al., 1999), they estimated ODP values for nPB ranging

1 Introduction

n-propyl bromide (chemical formula $CH_3CH_2CH_2Br$, also named 1-bromopropane and abbreviated below as nPB) has been proposed as a potential replacement for compounds that have been used as solvents for many years including



Correspondence to: D. J. Wuebbles
(wuebbles@atmos.uiuc.edu)

from 0.0002 (for summer emissions at 56° N to 64° N) to as large as 0.06 for tropical emissions. Bridgeman et al. (2000) used the TOMCAT 3-D global chemical-transport model ($5.6^\circ \times 5.6^\circ \times 31$ levels) to evaluate the atmospheric lifetime and amount of nPB reaching the stratosphere as a function of location and season of emission. They calculated ODPs empirically, based on the derived lifetime and the fraction of nPB emissions reaching the stratosphere. For nPB, they derived ODPs that vary from 0.0033 for emissions from Europe to 0.0109 for emissions from Indonesia. The study assumed that any bromine atoms released in the troposphere would be removed by rainout with 100% efficiency before reaching the stratosphere. Thus, their derived ODPs only considered the direct nPB reaching the stratosphere.

Wuebbles et al. (2001) used the MOZART-2 3-D chemical-transport model ($5^\circ \times 5^\circ \times 34$ levels) in combination with studies using their 2-D model in examining the potential effects of nPB on stratospheric ozone and the corresponding set of ODPs for nPB, and attempted to quantify degradation product effects. With available information, bromoacetone was the intermediate degradation product which had the largest concentration and had a local photochemical lifetime more than 1 day. For nPB emitted equally over global land masses north of 60° S and south of 70° N, their 3-D model calculations suggest that about 0.44% of the emitted bromine enters the stratosphere. About 33% of the bromine reaching the stratosphere results from the direct transport of nPB to the stratosphere and about 19% due to the transport of bromoacetone to the stratosphere. The rest (48%) is from transport of inorganic bromine.

These studies indicated the potential importance of considering both pathways in deriving the halogen reaching the stratosphere and the resulting ODPs for VSL gases like nPB. However, measurements by Burkholder et al. (2002) suggest that the local lifetime of bromoacetone is a few hours rather than at least one day as assumed in Wuebbles et al. (2001). Ko et al. (2003) modified the Wuebbles et al. (2001) ODP values to account for this overestimate of the bromoacetone atmospheric lifetime. For nPB emissions at midlatitudes (North America, Europe, and Asia – roughly 30° N to 60° N), the modified ODPs range from 0.017 to 0.026.

Since the earlier studies were published, much more has been learned about the oxidation of nPB and the chemistry of nPB degradation products in the atmosphere. In addition to the Burkholder et al. (2002) study of bromoacetone, there have been new measurements and analyses of nPB reaction with OH (Gilles et al., 2002; Martínez-Avilés et al., 2008a) and of the specific reactions and reaction rates in the reaction mechanism after initiation (Martínez-Avilés et al., 2008b). We have incorporated these new findings into the reaction mechanism used in the modeling studies.

2 Methodology

The ODP concept (Wuebbles, 1981, 1983) arose as a means of determining the relative ability of a chemical to destroy stratospheric ozone. ODPs are an integral part of national and international considerations on ozone-protection policy, including the Montreal Protocol and its Amendments and the US Clean Air Act. ODPs provide an important and relatively straightforward way of analyzing the potential for a new chemical to affect ozone relative to the CFCs, Halons, and other replacement compounds. ODPs are currently determined by two different means: calculations from chemistry-transport models of the global atmosphere (CTMs), and calculations using a semi-empirical approach (Solomon et al., 1992).

An ODP is a relative measure of the expected cumulative effect on stratospheric ozone per unit mass emission of a gas compared with the expected effect from the same mass emission of CFC-11. Therefore, it is defined as the change in total ozone per unit mass emission of the gas, relative to the change in total ozone per unit mass emission of CFC-11. ODP by itself does not, however, indicate the actual emissions or the amount of ozone depletion that has occurred or might occur.

Traditionally, zonally averaged 2-D CTMs were used for calculating the ozone response in the stratosphere. Past ODP evaluations were conducted for chemicals with atmospheric lifetimes more than approximately one year. Such long-lived gases mixed well throughout the troposphere after surface release, and large fractions of the surface emissions reached the stratosphere. However, many of the compounds now being considered either for new applications or as replacements for substances controlled under the Montreal Protocol are designed to be short-lived, on the order of days to a few months, so as to reduce the impacts on ozone and climate. These short-lived replacement gases still can be vertically transported into the lower stratosphere by intense convection in the tropical troposphere, which is relevant to possible effects on ozone from these compounds. Thus, the ODP of a VSL species depends upon its distribution in the atmosphere and the location of its source (Wuebbles and Ko, 1999; Wuebbles et al., 2001; Ko et al., 2003). Unfortunately, the determination of ODPs for these gases is not straightforward because their short atmospheric lifetimes leave them poorly mixed in the troposphere. Because they lose halogens in the troposphere, VSLs can also cause O₃ depletion in the troposphere, where both O₃ precursors and loss processes are highly variable with longitude and thus not easily treated in the zonal average. Also, the traditional 2-D model analysis of ODPs is not sufficiently accurate to calculate the integrated amount of the halogenated VSL source and reaction product gases in the troposphere that enter the stratosphere (Wuebbles et al., 2001).

3-D CTMs representing the complete troposphere and stratosphere are necessary for evaluating the halogen loading

and ozone depletion for VSL species. Thus the definition of ODPs has been revised for VSL compounds (Wuebbles et al., 2001; Ko et al., 2003). The new ODP definition for the VSL compounds accounts for the variation that can occur in the ODP as a function of where and when the compound is used and emitted. The most important factor in evaluating the ODP of VSL compounds is shown to be latitude distribution of the surface emissions because halogens from source gases emitted at higher latitudes are less likely to reach the stratosphere than from source gases emitted in the tropics (Bridgeman et al., 2000; Olsen et al., 2000; Wuebbles et al., 2001). 3-D CTMs, which include a much more comprehensive treatment of transport than available for 2-D models, are now preferable for calculations of ODPs both for the longer-lived gases and for the very short-lived candidate replacement compounds. These models can not only determine the amount of the substance reaching the stratosphere directly but can also follow the processes affecting the reaction products. Studies suggest that most chlorine or bromine that reaches the stratosphere from VSL substances is transported there in reaction products (Wuebbles et al., 2001; Ko et al., 2003; Law et al., 2007).

2.1 Atmospheric model

This study uses the current-generation 3-D model of global atmospheric chemistry and physics developed by the National Center for Atmospheric Research (NCAR) called the Model for Ozone And Related Tracers version 3.1 (MOZART-3). MOZART-3 has been used to explicitly calculate the impact of halocarbon emissions added at the Earth surface into the atmosphere on ozone depletion. The MOZART-3 CTM includes a complete representation of tropospheric, stratospheric, and upper atmospheric processes (Kinnison et al., 2007). It incorporates a full stratosphere, including the chemistry of chlorine species (Cl_y) and bromine species (Br_y) important in stratospheric ozone calculation as well as updated hydrogen, nitrogen, and hydrocarbon oxidation chemistry relevant to stratospheric and tropospheric chemistry included in the lower atmospheric version, MOZART-2 (Horowitz et al., 2003). The water-soluble Cl_y species HCl, HOCl, and ClONO_2 and the water-soluble Br_y species HBr, HOBr, and BrONO_2 are removed from the MOZART-3 atmosphere by tropospheric wet deposition with the same rate constant as used for HNO_3 . Representations of relevant heterogeneous and physical processes for winter/spring polar vortex related to ozone destruction are also fully included in the model. This model has been evaluated extensively via comparisons with measurements of atmospheric trace gases from satellite data and measurement campaigns (e.g., Wei et al., 2003; Pan et al., 2007).

Chemical reaction-rate constants and photochemical data for reactions other than those of nPB, its degradation product bromoacetone, TCE, and PCE follow the recommendations of Sander et al. (2003). The MOZART-3 CTM is driven

by meteorology fields derived from the Whole Atmosphere Community Climate Model (WACCM) version 1b (Sassi et al., 2004). MOZART-3 thus has a 2.8° resolution in latitude and longitude and a hybrid sigma-pressure vertical coordinate including 66 layers from the surface to 5.1×10^{-6} hPa (approximately 140 km).

2.2 Atmospheric oxidation of nPB, TCE, and PCE

In this study, the initiating reaction of nPB with OH radical is based on the rate constants from Gilles et al. (2002):

$$k(\text{nPB} + \text{OH} \rightarrow \text{CH}_3\text{CH}_2\text{CHBr} + \text{H}_2\text{O}) = 1.44 \times 10^{-12} \exp(-450/T) \quad (1a)$$

$$k(\text{nPB} + \text{OH} \rightarrow \text{CH}_3\text{CHCH}_2\text{Br} + \text{H}_2\text{O}) = 2.54 \times 10^{-18} T^2 \exp(265/T) \quad (1b)$$

$$k(\text{nPB} + \text{OH} \rightarrow \text{CH}_2\text{CH}_2\text{CH}_2\text{Br} + \text{H}_2\text{O}) = 2.89 \times 10^{-12} \exp(-957/T) \quad (1c)$$

where the temperature T is in K and rate constants k are in $\text{cm}^3 \text{molec}^{-1} \text{s}^{-1}$. Analysis of the subsequent reactions by the radicals produced by OH reaction with nPB from Martínez-Avilés et al. (2008a, b) suggests that the only organobromine degradation product to have a significant lifetime against further reaction and to be produced in a significant fraction of the nPB input to the atmosphere is bromoacetone, $\text{CH}_3\text{C}(\text{O})\text{CH}_2\text{Br}$ (abbreviated as BrAc below). The absorption cross sections and photolysis quantum yields for BrAc of Burkholder et al. (2002) are used in this study.

The chloroalkenes PCE and TCE are expected to react in the atmosphere primarily by addition of OH radical to the C=C double bond. The rate constants for these addition reactions are taken as the high-pressure limit expressions from the JPL 2006 photochemical data evaluation by Sander et al. (2006):

$$k(\text{PCE} + \text{OH}) = 4.7 \times 10^{-12} \exp(-990/T) \quad (2)$$

$$k(\text{TCE} + \text{OH}) = 8.00 \times 10^{-13} \exp(300/T) \quad (3)$$

where the temperature T is in K and rate constants k are in $\text{cm}^3 \text{molec}^{-1} \text{s}^{-1}$. The PCE + OH rate constant expression in the IUPAC photochemical data evaluation (Atkinson et al., 2008) agrees with that in Sander et al. (2006) to within 6% for T of 200 to 300 K. The TCE + OH rate constant expression from Atkinson et al. (2008) agrees with that given by Sander et al. (2006) for T of 240 to 300 K; while the IUPAC values are more than 10% higher for T below 240 K, this is not expected to produce much difference in the TCE lifetime because limited amounts of TCE reach colder temperatures in the upper troposphere (see Fig. 4a below).

Radicals produced by OH addition to PCE or TCE then are likely to add O_2 , then react with NO to produce chlorinated hydroxyalkoxy radicals. In this study, we have approximated the subsequent chemistry as release of all chlorine within one time step. The chlorinated hydroxyalkoxy radicals are likely to produce phosgene (CCl_2O) for both TCE and PCE and chloroformaldehyde (CHClO) for TCE, which react further by both OH and photolysis under daytime conditions suitable for oxidation of PCE or TCE.

2.3 ODP from MOZART-3

We evaluate the atmospheric lifetimes and Ozone Depletion Potentials (ODPs) for the three compounds in the MOZART-3 CTM. A steady-state background atmosphere corresponding to the year 2000 was derived assuming fixed surface mixing ratios for long-lived gases (such as nitrous oxide, methane, chlorofluorocarbons, and other halogenated source gases) and fixed emissions for short-lived gases. An ODP calculation in an atmospheric CTM requires two perturbation runs for comparison to the CTM reference atmosphere: a CFC-11 (CFC1₃, also known as trichlorofluoromethane) run and a run incorporating the compound under evaluation. The perturbations of CFC-11 or of the compound under evaluation are selected to give similar decreases in globally-averaged O₃ that are less than one percent. Emissions of nPB, TCE, and PCE were assumed to occur entirely on land at midlatitudes in the Northern Hemisphere, from 30° N to 60° N (which corresponds to where the vast majority of current or future emissions are also expected to occur). In addition, for nPB the atmospheric lifetime and ODP was determined assuming emissions evenly distributed over land globally, from 60° S to 70° N. Each CTM perturbation is run to steady state, after which the change in total O₃ burden in the atmosphere and the flux of the substance causing the ozone change are calculated for the CFC-11 run and for the run of the compound under study. The ODPs calculated in this study use the global total O₃ burden loss because O₃ absorbs harmful solar ultraviolet radiation whether or not it is located in the stratosphere; we do not attempt to separate ODP into tropospheric and stratospheric components, though we do note the fraction of O₃ burden loss that occurs in the troposphere for nPB, TCE, and PCE below.

The CFC-11 perturbation used in MOZART-3 for this study is an increase of the fixed surface mixing ratio by 80 ppt. This perturbation required 10 years of model time to reach a steady state global O₃ burden loss of -0.594%. The change in CFC-11 loss rate due to that perturbation is 0.0335 Tg yr⁻¹, which equals the increase of CFC-11 flux at the surface in the model. nPB, TCE, and PCE fluxes added to the model in this study are chosen to provide global O₃ burden losses less than 1% and similar, ideally to within a factor of two, to 0.594% as obtained for CFC-11; these fluxes are not intended to represent historic, current, or likely future emissions of any of these compounds. We ran all of the model calculations on the same compute platform in order to avoid possible portability issues.

3 Results

This study considers three VSLs with a variety of surface fluxes introduced into the MOZART-3 CTM and a significant range of atmospheric lifetimes. In order to enable comparison among these results, we sought to present them on a

common basis in the figures that follow. ODP is a measure of relative decrease in global O₃ burden, so that we select 1% decrease in global O₃ burden arising from each VSL scenario as that common basis. Because the O₃ burden loss is expected to be proportional to the flux of each VSL and to the increase in Cl_y or Br_y resulting from addition of that VSL to MOZART-3, we divide each model output (annual- and zonal-average VSL mixing ratios, Cl_y or Br_y mixing ratio perturbations, percent changes in O₃ mixing ratio, and BrAc mixing ratios for nPB as the VSL) presented in the figures below by the total global O₃ burden loss averaged over the year for that VSL as calculated in MOZART-3. We refer to these values from MOZART-3 divided by global O₃ burden loss by the phrase “per 1% O₃ burden loss” in this discussion.

3.1 nPB

The nPB flux used to illustrate industrial use is a rate of 8.64×10^8 molecules cm⁻² s⁻¹ from all land at latitudes 30° N to 60° N, for a total nPB emission rate of 2.48 Tg yr⁻¹. That flux required 7 model years to reach steady state and resulted in a global O₃ burden decrease of 0.214%, of which 0.123% was above the tropopause. Figure 1 shows the derived annual and zonal average distribution of nPB, the corresponding distribution of BrAc and change in Br_y, and the resulting change in the distribution of tropospheric and stratospheric O₃ divided by the 0.214% decrease in global O₃ burden. In each part of this and subsequent figures, the dotted white line indicates the average pressure level of the MOZART tropopause. The nPB mixing ratio per 1% O₃ burden loss in Fig. 1a peaks at the emissions region with little reaching the Southern Hemisphere, and no more than 10 ppt, or 2% of the 349 ppt maximum mixing ratio, reaches the tropopause. The potential product gas injection BrAc mixing ratio per 1% O₃ burden loss peaks at 2.1 ppt near the surface in Fig. 1b, and its maximum mixing ratio at the tropopause is less than 0.02 ppt. However, the Br_y perturbation per 1% O₃ burden loss in the Northern Hemisphere exceeds 2.5 ppt for much of the lower stratosphere, as shown in Fig. 1c. The stratospheric Br_y perturbation in excess of nPB mixing ratios at the tropopause suggests that Br_y in the stratosphere is increased by some means besides the nPB that reaches the stratosphere. The minimum Br_y perturbation in the tropics suggests that the remainder of the stratospheric Br_y perturbation could be due to a decrease in the net Br_y transport from the stratosphere to the troposphere caused by the Br_y increase in the Northern troposphere. nPB increases Br_y by at least 1 ppt everywhere in and above the MOZART stratosphere, and the O₃ perturbation percentage per 1% O₃ burden loss in Fig. 1d includes both a peak in the Northern polar upper troposphere and a secondary maximum near the South Polar tropopause with minimal O₃ loss in the tropics around 10 hPa. The atmospheric chemical lifetime obtained

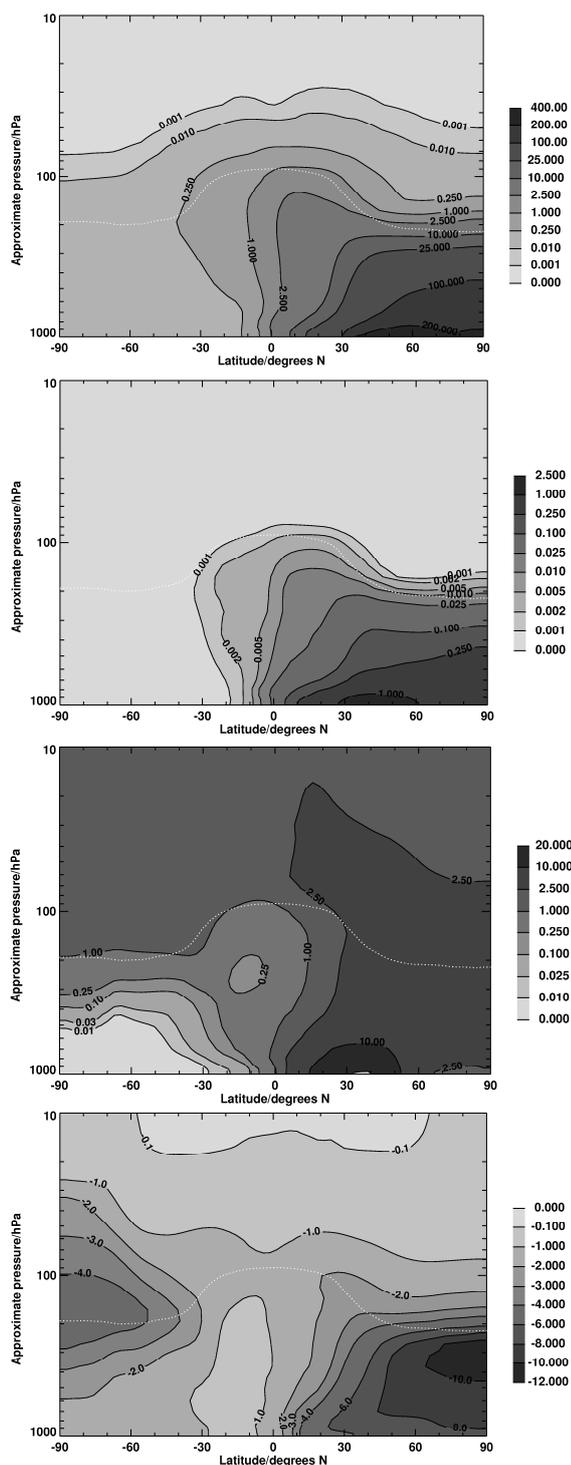


Fig. 1. Annual- and zonal-average mixing ratio changes calculated in MOZART-3 for n-propyl bromide (nPB) emissions evenly distributed on land surfaces from 30° N to 60° N totaling 2.48 Tg yr^{-1} , divided by 0.214 to correspond to 1% decrease in global total ozone (O_3) burden. (a) nPB in ppt; (b) bromoacetone (BrAc) in ppt; (c) inorganic bromine species (Br_y) in ppt; (d) O_3 in percent. In this and all following figures, the dotted white line represents the annual- and zonal-average tropopause.

for nPB is 24.7 days, comparable to the value reported in Ko et al. (2003), and the chemical lifetime for BrAc is 5.4 h, consistent with the estimate of Burkholder et al. (2002). The ODP obtained for nPB is 0.0049, smaller than the value reported previously. This ODP is lower by a factor of more than three compared to the most comparable value reported in Tables 2–12 of Ko et al. (2003), 0.017 for emissions from North America, Europe, and Asia with the correction for reduced BrAc using maximum rainout of Br_y species in Wuebbles et al. (2001). Atmospheric lifetimes and ODPs for nPB, TCE, and PCE obtained in this study are summarized in Table 1.

A study with nPB emissions of $2.90 \times 10^8 \text{ molecules cm}^{-2} \text{ s}^{-1}$ evenly distributed over land surfaces from 60° S to 70° N, which totals 2.49 Tg yr^{-1} and is comparable to the "global" case in Wuebbles et al. (2001), was also conducted in MOZART-3; this perturbation also required 7 model years to reach steady state with a global O_3 column change of -0.470% . The nPB annual and zonal average mixing ratios divided by 0.470 to represent 1% global O_3 burden loss, shown in Fig. 2a, show that penetration to the tropopause is more symmetric than for 30°–60° N emissions, and between 1 and 2.5 ppt of nPB reaches the tropopause at all latitudes north of 25° S. While the BrAc distribution per 1% O_3 burden loss in Fig. 2b is also more symmetric than that for 30°–60° N nPB emissions, the peak mixing ratio is 0.43 ppt compared to a maximum nPB mixing ratio of 110 ppt in Fig. 2a, and the maximum tropopause BrAc mixing ratio is less than 0.005 ppt. The Br_y perturbation per 1% O_3 burden loss from these nPB emissions, shown in Fig. 2c, is more than 2.5 ppt throughout much of the stratosphere up to the 10 hPa pressure layer, but the Br_y perturbation at the tropopause is smaller for 30° S to 30° N than for surrounding latitudes outside that range. As with Fig. 1a–c, the fact that stratospheric Br_y increases more than the maximum stratospheric nPB mixing ratio indicates that in addition to the component from nPB reaching the stratosphere, stratospheric Br_y from nPB in MOZART-3 could be increasing in part because of decreased net Br_y transport from the stratosphere to the troposphere resulting from increased Br_y concentrations in the extratropical troposphere. The O_3 percent loss profile per 1% O_3 burden loss from these nPB emissions shown in Fig. 2d again peaks at either pole, but the South Polar tropopause is the maximum loss and the reduction in tropical tropospheric O_3 is greater than for 30°–60° N emissions in Fig. 1d. For the global emissions case, the derived atmospheric lifetime of nPB is 19.6 days, nearly the same as in Wuebbles et al. (2001), and the resulting ODP is 0.011, more than a factor of two lower than the global high wet-deposition ODP reported in Wuebbles et al. (2001) or the 0.027 value given in Ko et al. (2003) Table 2–12 for that Wuebbles et al. (2001) result after correction for the revised BrAc lifetime. The comparisons of current ODPs with the values reported in Ko et al. (2003) after BrAc lifetime correction indicate that O_3 in MOZART-3 is less sensitive to Br_y from nPB than was the

Table 1. Derived atmospheric lifetimes and ODPs.

Gas	Emissions Latitudes	Lifetime, days	ODP
nPB	30° N–60° N	24.7	0.0049
nPB	60° S–70° N	19.6	0.011
TCE	30° N–60° N	13.0	0.00037
PCE	30° N–60° N	111	0.0050

Table 2. Summary of lifetimes and ozone changes in MOZART-3 for two total surface fluxes of trichloroethylene (TCE) from 30° N to 60° N land surfaces.

TCE flux, Tg yr ⁻¹	Lifetime, days	O ₃ column change, %	ODP
12.9	13.0	–0.0851	0.00037
51.7	13.8	–0.338	0.00037

case for the combination of MOZART-2 with the UIUC 2-D model as used in Wuebbles et al. (2001).

3.2 PCE and TCE

The MOZART-3 CTM with added surface flux for land from 30° N to 60° N is run to near-steady-state for TCE and for PCE as with nPB fluxes. After 9 years model time to reach steady state, a PCE surface flux of 3.91 Tg yr⁻¹ in MOZART-3 results in a global O₃ column change of –0.344%. The annual and zonal average mixing ratio distribution due to PCE divided by 0.344 to represent 1% global O₃ burden loss is shown in Fig. 3a. PCE reaches the tropical tropopause at a markedly higher fraction of the peak mixing ratio than for nPB (compare Fig. 1a). The tropopause PCE mixing ratios, which range from 100 ppt at the North Pole to less than 50 ppt poleward of 40° S, are consistent with the stratospheric Cl_y perturbation per 1% O₃ burden loss of up to 0.35 ppb in Fig. 3b because each PCE molecule carries four chlorine atoms and because the small values of tropospheric Cl_y perturbation form a similar pattern to those from longer-lived chlorocarbons such as HCFC-123. PCE seems likely to transport Cl_y to the stratosphere as PCE by a combination of tropical and extra-tropical transport, unlike the case for Br_y from nPB. Figure 3c shows that O₃ loss from PCE per 1% O₃ burden loss occurs mostly at polar latitudes with a maximum loss in the South Polar lower stratosphere. The PCE atmospheric lifetime obtained is 111 days, and the ODP calculated for PCE in MOZART is 0.0050, among the lowest of values obtained for any chlorocarbon so far.

TCE was run at two surface fluxes for which the MOZART results are summarized in Table 2. The initial run using 51.7 Tg yr⁻¹ with 7 model years to reach steady state produced a global O₃ column decrease of 0.338% with an atmospheric lifetime of 13.8 days. Figure 4a shows the annual and

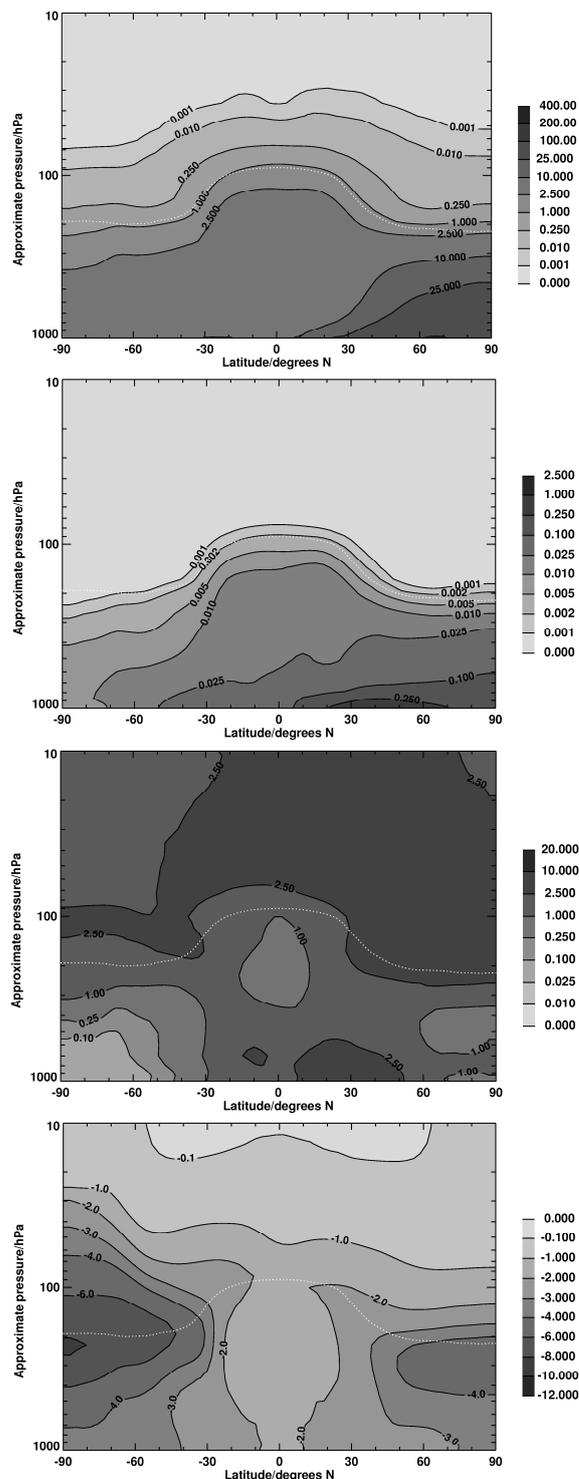


Fig. 2. Annual- and zonal-average mixing ratio changes calculated in MOZART-3 for nPB emissions evenly distributed on land surfaces from 60° S to 70° N totaling 2.49 Tg yr⁻¹, divided by 0.470 to correspond to 1% decrease in global total O₃ burden. (a) nPB in ppt; (b) BrAc in ppt; (c) Br_y in ppt; (d) O₃ in percent.

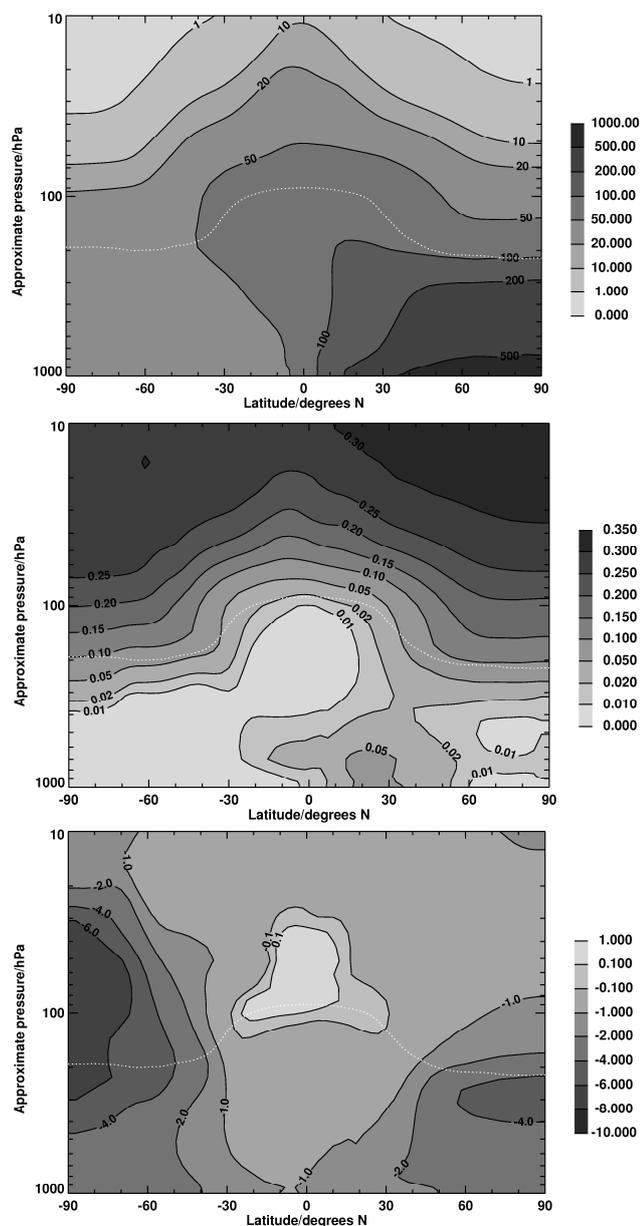


Fig. 3. Annual- and zonal-average mixing ratio changes calculated in MOZART-3 for perchloroethylene (PCE) emissions evenly distributed on land surfaces from 30° N to 60° N totaling 3.91 Tg yr⁻¹, divided by 0.344 to correspond to 1% decrease in global total O₃ burden. (a) PCE in ppt; (b) Cl_y in ppb; (c) O₃ in percent.

zonal average TCE distribution divided by 0.338 to represent a 1% global O₃ burden loss. The maximum TCE mixing ratio per 1% O₃ burden loss is 5750 ppt in the emissions region, and intact TCE transport to the stratosphere is limited to perhaps 0.1% even at the North Pole. The Cl_y perturbations (ppb) per 1% O₃ burden loss in Fig. 4b are concentrated in the emission region, with mixing of Cl_y from TCE along the extratropical Northern upper troposphere and into the stratosphere seemingly limited by rainout removal at North Polar

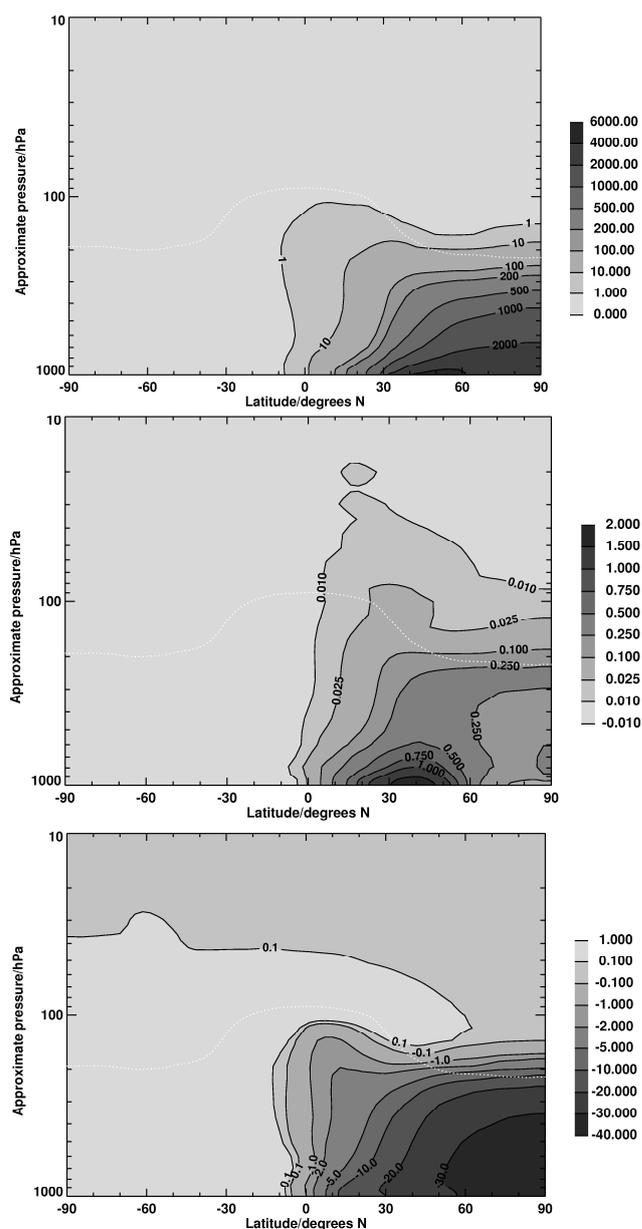


Fig. 4. Annual- and zonal-average mixing ratio changes calculated in MOZART-3 for trichloroethylene (TCE) emissions evenly distributed on land surfaces from 30° N to 60° N totaling 51.7 Tg yr⁻¹, divided by 0.338 to correspond to 1% decrease in global total O₃ burden. (a) TCE in ppt; (b) inorganic chlorine species (Cl_y) in ppb; (c) O₃ in percent.

latitudes below the 300 hPa pressure layer and effectively no change in Cl_y throughout most of the stratosphere and the Southern troposphere. The O₃ losses resulting from TCE per 1% O₃ burden loss shown in Fig. 4c are mostly constrained to the Northern troposphere and range up to 40% reduction. Because this large TCE flux results in such a large change of O₃ at Northern latitudes, distortion of the OH field in MOZART resulting in an overestimation of TCE lifetime

is a possible concern. Therefore, we also ran with one-fourth that TCE flux, or 12.9 Tg yr^{-1} . After 8 model years to reach steady state (due to an operational problem in model year 7 of this run), the O_3 column perturbation was one-fourth that of the 51.7 Tg yr^{-1} run, and the TCE lifetime reported was 13.0 days, consistent with a minimal OH field perturbation from TCE, and we report the lifetime from the 12.9 Tg yr^{-1} run in Table 1. The TCE ODP, 0.00037, is unaffected to within 1% by the change of surface flux.

The fluxes of TCE and PCE used in this study were selected in order to produce numerically significant O_3 loss in MOZART-3, so that these fluxes are markedly larger than the global production of these compounds. The ODP results suggest that historic annual releases of these two chlorocarbons to the atmosphere, which as of 1999 totaled 0.146 Tg for TCE and 0.387 Tg for PCE in the Reactive Chlorine Emissions Inventory (McCulloch et al., 1999; Keene et al., 1999), should have less than 0.1% effects on global total ozone. Global emissions of TCE and PCE have been decreasing since 1999, and emissions of these compounds will likely continue to decrease in the future, so that these compounds likely produce negligible ozone effects.

Our MOZART-3 analyses of TCE and PCE have not considered possible degradation chemistry after the initial addition of OH, instead treating these compounds as if all Cl atoms were released within a model time step. TCE and PCE atmospheric degradation schemes have not been comprehensively evaluated by experiment, but a recent study (Nolan et al., 2006) indicates that some 70% of TCE and PCE reacting with OH produces phosgene (COCl_2) under atmospheric conditions. COCl_2 has negligible reactivity with gas-phase species in the troposphere and a 1.85 year lifetime against photolysis in the stratosphere (Kindler et al., 1995), suggesting that it could transport significant Cl from TCE or PCE to the stratosphere. However, COCl_2 hydrolyzes in liquid water (Sander et al., 2006, and references therein) so that this oxidation product could reduce total gas-phase Cl perturbations resulting from TCE or PCE. An additional TCE and PCE oxidation product likely to be produced by the mechanisms presented in Nolan et al. (2006) is hydrogen chloride (HCl), which is soluble in liquid water and readily adsorbed on ice. While HCl yield is not quantified in the Nolan et al. study, HCl removal by wet deposition would reduce the total Cl input to the atmosphere from TCE or PCE emissions. If this effect applies beyond the existing MOZART-3 wet deposition of HCl, HOCl, and ClONO_2 produced in the model troposphere by Cl released from these compounds, TCE and PCE ODPs reported in this study would be overestimated because of the possible production of HCl and COCl_2 by tropospheric degradation, each of which is subject to wet deposition.

3.3 ODP uncertainties

Our study incorporates the VSLs nPB, PCE, and TCE into a global 3-D CTM that fully represents the upper atmosphere

and includes a more comprehensive representation of atmospheric chemistry and transport relevant to the expected effects of these VSLs than in past studies of nPB. However, the nPB ODPs reported in this study are considerably lower than those in Wuebbles et al. (2001), and the decrease in BrAc lifetime is not likely to have caused such a large reduction in ODPs. The apparent lower sensitivity of O_3 in MOZART-3 to Br_y from nPB compared to that of O_3 in the combination of MOZART-2 with our two-dimensional chemical-transport model (2-D CTM) could arise from several differences in these models, including:

- MOZART-3 uses the JPL02 chemistry recommendation (Sander et al., 2003); MOZART-2 and our 2-D CTM used the prior version.
- Transport in the full-atmosphere MOZART-3 has been revised compared to the tropospheric model, MOZART-2, in tropospheric and stratospheric circulation and in exchange across the tropopause.
- Faster rainfall rates may be obtained in MOZART-3 than in MOZART-2 or the 2-D CTM, leading to faster Br_y wet deposition.

We did not save information regarding Br_y wet deposition in our MOZART-3 runs, so that we cannot definitely identify the possible contribution of faster rainout to the lower ODPs obtained for nPB in this study with our existing output.

The lifting speed of VSLs such as nPB, TCE, and PCE by the MOZART-3 convection process is an important component of their predicted effects on O_3 . The convective lifting speed is difficult to evaluate against measurements because wind information must be obtained at considerable detail that is unwieldy to use in a model. However, the convection parameterization in WACCM 1b must be reasonable in order to allow it to obtain a good representation of temperature and of specific humidity. Because MOZART-3 uses similar convection, we believe it to be reasonable as well with the caveat that it has not been fully evaluated.

MOZART-3, as with other atmospheric models of its generation, uses a somewhat coarse vertical resolution near the tropopause of just under 1 km. This coarse vertical resolution causes MOZART-3 to underestimate steep gradients in the tropopause region (Pan et al., 2007) and could cause an overestimation of the amount of VSLs that reaches the stratosphere. If this is taking place for nPB, TCE, or PCE, the ODPs reported in this study would be overestimated.

Transport in newer models, including climate-chemistry models such as version 3 of WACCM, is being evaluated against measurements in more detail than was applied to MOZART-3 or WACCM 1b (Gettelman et al., 2010; Neu et al., 2010). In these tests, WACCM 3 was one of the best of the climate-chemistry models in representing transport. While WACCM 3 does include updated processes compared to WACCM 1b, the earlier version of WACCM, and thus

MOZART-3 when driven with its meteorology, is likely to have represented atmospheric processing of VSLs as well as any model of their generation. Transport is nonetheless likely to be one of the most significant components of uncertainty in the ODPs derived in this study, and those values should be taken as uncertain by at least 10% to 25%.

4 Summary

We have calculated atmospheric lifetimes and Ozone Depletion Potentials (ODPs) for the currently-used compounds trichloroethylene (TCE) and perchloroethylene (PCE) and the proposed replacement compound n-propyl bromide (nPB) in the MOZART-3 chemical-transport model of the global atmosphere. Unlike previous studies of nPB in particular, MOZART-3 represents the upper atmosphere fully and includes detailed chemical and transport processes, which leads to a more comprehensive prediction of the effects on ozone to be expected from these gases. For emissions of nPB from all land surfaces between 30° N and 60° N, we obtain a lifetime of 24.7 days and an ODP of 0.0049. Emissions of nPB from all land between 60° S and 70° N result in a lifetime of 19.6 days and an ODP of 0.011. These are among the lowest ODPs obtained for a compound containing bromine in our studies so far, and the nPB ODP is lower than obtained in past studies. For PCE emissions from all land surfaces between 30° N and 60° N, the atmospheric lifetime is 111 days and the ODP is 0.0050, and for TCE emissions from land from 30° N to 60° N, the lifetime is 13.0 days and the ODP is 0.00037. Neither nPB nor TCE are predicted to reach the stratosphere in large quantity to release bromine or chlorine directly, but significant PCE does reach the stratosphere intact to release chlorine.

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